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the N_2O_4 gas decomposes to NO_2 gas and the reaction reaches equilibrium according to the following balanced equation $\text{N}_2\text{O}_4(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g}) \quad \Delta H > 0$ the graph below shows how the concentrations of the two gases change as a result of changes made to the reaction conditions $\text{N}_2\text{O}_4(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$

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2 2 vf vi 02 vi 2 2 9 8 0 9 1 vi upwards opwaarts 4 criteria for graph kriteria vir grafiek marks punte 1 first part of the graph starts at v 2 at t 0 s and extends until v 6 27 1 at t 0 44 s 1 eerste deel van die grafiek begin by v 2 by t 0 s en 1 verleng tot v 6 27 by t 0 44 s graph is discontinuous and

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two blocks of masses 20 kg and 5 kg respectively are connected by a light inextensible string p a second light inextensible string q attached to the 5 kg block runs over a light frictionless pulley

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