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the periodic table organizes elements into groups and periods based on their chemical and physical properties elements in the same group share similar characteristics like reactivity the table is divided into metals nonmetals and metalloids each with distinct properties

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learn how the periodic table of the elements was created organized and classified find out the properties and trends of metals nonmetals and metalloids

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periodic table in chemistry the organized array of all the chemical elements in order of increasing atomic number when the elements are thus arranged there is a recurring pattern called the periodic law in their properties in which elements in the same column group have similar properties

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in this unit we ll cover some of chemistry s most fundamental topics including atoms isotopes ions and the periodic table we ll also explore different ways of representing atoms such as through bohr models and lewis diagrams

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the periodic table of elements puts all the known elements into groups with similar properties this makes it an important tool for chemists nanotechnologists and other scientists if you get to understand the periodic table and learn to use it you ll be able to predict how chemicals will behave

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the periodic table is an arrangement of the elements in order of increasing atomic number elements that exhibit similar chemistry appear in vertical columns called groups numbered 1 18 from left to right the seven horizontal rows are called periods

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scientists use the periodic table to classify chemical elements into different categories there are several different trends within the table that can help you understand how an element or group of elements will behave by itself and

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the periodic table of elements is a concise information dense catalog of all of the different sorts of atoms in the universe and it has a wealth of information to tell us if we can learn to read it

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the modern periodic table emphasizes the electronic structure of atoms the original periodic table emphasized the reactivity of the elements in this chapter we will learn about the connection between the two and how the electronic structure is related to the macroscopic properties of the elements

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