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quantum mechanics is often the only theory that can reveal the individual behaviors of the subatomic particles that make up all forms of matter electrons protons neutrons photons and others solid state physics and materials science are dependent upon quantum mechanics

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quantum mechanics is the subfield of physics that describes this bizarre behavior of microscopic particles atoms electrons photons and almost everything else in the molecular and

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quantum mechanics science dealing with the behavior of matter and light on the atomic and subatomic scale it attempts to describe and account for the properties of molecules and atoms and their constituents electrons protons neutrons and other more esoteric particles such as quarks and gluons

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quantum mechanics is the study of matter and its interactions with energy on the scale of atomic and subatomic particles by contrast classical physics explains matter and energy only on a scale familiar to human experience including the behavior of astronomical bodies such as the moon

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quantum mechanics qm is a branch of physics which describes physical systems so that properties like the energy or angular momentum are discrete quantities that are multiples of a smallest unit or quantum a famous physicist named erwin schrödinger made an example of quantum mechanics often called schrödinger s cat

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my best answer so far is that quantum mechanics is the study of how small particles move and interact but that s an incomplete answer since quantum effects can be important on galactic scales too

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quantum mechanics is the field of physics that explains how extremely small objects simultaneously have the characteristics of both particles tiny pieces of matter and waves a disturbance or variation that transfers energy physicists call this the wave particle duality wave particle duality and quanta

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key points louis de broglie proposed that all particles could be treated as matter waves with a wavelength λ given by the following equation $\lambda = h / m \cdot v$ erwin schrödinger proposed the quantum mechanical model of the atom which treats electrons as matter waves schrödinger s equation $\hbar \psi = E \psi$

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all are based upon quantum theory i e quantum mechanics and quantum field theory and relativity which together form the theoretical foundations of modern physics many physical quantities whose classical counterparts vary continuously over a range of possible values are in quantum theory constrained

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what you ll find below is a concise clear eyed guide to the known particles and forces from electrons quarks and neutrinos to photons and the higgs boson as well as the quantum laws and

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metrologists have measured planck s constant to be $6.626070150 \times 10^{-34}$ kg m² s⁻¹ a value that is correct to within 10 parts per billion in our explore physics series we delve into the quirky universe of quantum mechanics where objects behave as both particles and waves

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it describes the outcomes of experiments that showed that light and matter had the properties of particles or waves depending on how they were measured today we understand that these different forms of energy are actually neither particle nor wave they are distinct quantum objects that we cannot easily conceptualize

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in quantum mechanics particles have wavelike properties and a particular wave equation the schrodinger equation governs how these waves behave the schrodinger equation is different in a few ways from the other wave equations we've seen in this book

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in quantum mechanics the state of a physical system is represented by a wave function in born's interpretation the square of the particle's wave function represents the probability density of finding the particle around a specific location in space

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all of nature springs from a handful of components the fundamental particles that interact with one another in only a few different ways in the 1970s physicists developed a set of equations describing these particles and interactions together the equations formed a succinct theory now known as the standard model of particle physics

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quantum mechanics revealed to its discoverers in the 1920s that photons and other quantum objects are best described not as particles or waves but by abstract wave functions evolving mathematical functions that indicate a particle's probability of having various properties

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