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solution in chemistry a homogenous mixture of two or more substances in relative amounts that can be varied continuously up to what is called the limit of solubility the term solution is commonly applied to the liquid state of matter but solutions of gases and solids are possible we call the major component the solvent and the minor component s the solute s the most familiar solutions are aqueous solutions in which water is the solvent for example in a solution of the sugar glucose in water glucose molecules are the solute and water molecules are the solvent in chemistry a solution is a special type of homogeneous mixture composed of two or more substances in such a mixture a solute is a substance dissolved in another substance known as a solvent a solution is a homogeneous mixture comprising smaller component s called solute s of small molecules or ions comparable in size to the molecules of a larger component called the solvent for example nacl dissolved in water is a solution the solute is almost uniformly distributed in the solvent making a homogeneous mixture the meaning of solution is an action or process of solving a problem how to use solution in a sentence the most common unit of concentration is molarity which is also the most useful for calculations involving the stoichiometry of reactions in solution the molarity m is defined as the number of moles of solute present in exactly 1 l of solution a solution is a homogeneous mixture of two or more substances a solution may exist in any phase a solution consists of a solute and a solvent the solute is the substance that is dissolved in the solvent the amount of solute that can be dissolved in solvent is called its solubility the act of solving a problem question etc the situation is approaching solution the state of being solved a problem capable of solution a particular instance or method of solving an explanation or answer the solution is as good as any other synonyms resolution key mathematics the process of determining the answer to a problem a solution is a homogeneous mixture of one or more solutes dissolved in a solvent solvent the substance in which a solute dissolves to produce a homogeneous mixture solute the substance that dissolves in a solvent to produce a homogeneous mixture note that the solvent is the substance that is present in the greatest amount a mixture in which one substance is dissolved in another especially a solid dissolved in a liquid an aqueous solution of salts in solution the warmer water becomes the less oxygen it can hold in solution dissolved in it smart vocabulary related words and phrases definitions of solution solute and solvent how molarity is used to quantify the concentration of solute and how to calculate molarity key points mixtures with uniform composition are called homogeneous mixtures or solutions mixtures with non uniform composition are heterogeneous mixtures this molarity calculator is a tool for converting the mass concentration of any solution to molar concentration or recalculating grams per ml to moles you can also calculate the mass of a substance needed to achieve a desired molarity this article will provide you with the molarity definition and the molarity formula introduction in aqueous solution an acid is defined as any species that increases the concentration of h^+ while a base increases the concentration of oh^- typical concentrations of these ions in solution can be very small and they also span a wide range the color of hydrangea flowers can vary depending on the ph of the soil a solution is a homogeneous mixture of two or more components in which the particle size is smaller than 1 nm common examples of solutions are sugar in water and salt in water solutions soda water

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