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the reaction rate for a given reaction is a crucial tool that enables us to calculate the specific order of a reaction the order of a reaction is important in that it enables us to classify specific chemical reactions easily and efficiently

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rate laws and reaction order the relation between the rate of a reaction and the concentrations of reactants is expressed by its rate law for example the rate of the gas phase decomposition of dinitrogen pentoxide  $2\text{N}_2\text{O}_5 \rightarrow 4\text{NO}_2 + \text{O}_2$

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the rate of a reaction also shorthand as rate of reaction has a simple definition that when uncovered has a lot of nuance to it the simple definition of the rate of reaction is how quickly a reaction produces products however this definition brings up some issues how do we measure how quickly a reaction occurs

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the rate of reaction is the change in the amount of a reactant or product per unit time reaction rates are therefore determined by measuring the time dependence of some property that can be related to reactant or product amounts

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the rate equation is  $\text{rate} = k[\text{P}]^2$  complete the table to show how changing the concentrations affects the rate initial rate  $\text{mol dm}^{-3} \text{ s}^{-1}$  change in concentration of reagents effect on rate new initial rate  $\text{mol dm}^{-3} \text{ s}^{-1}$

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a chemical reaction or process in which more energy is released than is required to break bonds in the initial reactants study with quizlet and memorize flashcards containing terms like reaction rate collision theory activated complex and more

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university of colorado department of physics what makes a reaction happen find out what affects the rate of a reaction do experiments collect data and then calculate rate coefficients play with different reactions concentrations and energy the reactions rates simulation dev

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definition the rate of a chemical reaction is a measure of how fast a reaction takes place it is defined as the change in concentration of a reactant or product per unit time for example in a reaction between magnesium and hydrochloric acid the rate could be measured in terms of the change in concentration of the hydrochloric acid per second

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consider the curves below point a forward reaction at the start of the reaction the rate of the forward reaction is faster the concentration of the reactants is greatest at the beginning the rate of forward reaction and the concentration of reactants decreases with time as the reaction proceeds

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## **a guide to rate of reaction Jun 29 2023**

what factors affect the rate of reactions in this lesson we go through the five factors that affect the rate of a reaction state of division of a solid surface area concentration pressure of a gas temperature and catalysts we show the effect of a change in each of these factors 2 how to measure reaction rates

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reaction rates 4 x conical flask y balance scales calcium carbonate hydrochloric acid calcium chloride water carbon dioxide  $\text{CaCO}_3$   $2\text{HCl}$   $\text{CaCl}_2$   $\text{H}_2\text{O}$   $\text{CO}_2$  carbon dioxide gas escapes stops liquid escaping

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