

## Free epub Stechiometria [PDF]

Stechiometria is a branch of chemistry that deals with the quantitative relationships between the elements in a chemical compound. It is based on the law of conservation of mass, which states that matter is neither created nor destroyed in a chemical reaction. The study of stoichiometry involves the calculation of the relative masses of the elements in a compound, the determination of the empirical formula of a compound, and the calculation of the amount of a substance that can be produced from a given amount of reactants.

The basic principle of stoichiometry is the conservation of mass. In a chemical reaction, the total mass of the reactants is equal to the total mass of the products. This principle is used to calculate the relative masses of the elements in a compound and to determine the empirical formula of a compound. The empirical formula is the simplest whole-number ratio of the atoms in a compound. It is determined by dividing the mass of each element in the compound by its atomic weight and then dividing the resulting numbers by the smallest number to get the relative number of atoms of each element.

Stoichiometry is also used to calculate the amount of a substance that can be produced from a given amount of reactants. This is done by using the balanced chemical equation for the reaction and the relative masses of the reactants and products. The amount of a substance that can be produced is limited by the amount of the reactant that is present in the smallest amount, which is called the limiting reactant.

Stoichiometry is a fundamental concept in chemistry and is used in many areas of science and industry. It is used to determine the composition of a compound, to calculate the amount of a substance that can be produced from a given amount of reactants, and to understand the quantitative relationships between the elements in a chemical compound.

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