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## **bases and salts classnotes ng**

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a base is a substance which will neutralize an acid to yield a salt and water only most oxide and hydroxide of metals are bases e g  $\text{Na}_2\text{O}$   $\text{K}_2\text{O}$   $\text{MgO}$   $\text{NaOH}$   $\text{KOH}$  e t c an alkali is a basic hydroxide which is soluble in water  $\text{NaOH}$   $\text{KOH}$   $\text{Ca(OH)}_2$

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learn the basics of acid base chemistry get the definitions and examples of acids and bases and learn about the neutralization reaction

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acid base reactions are essential in both biochemistry and industrial chemistry moreover many of the substances we encounter in our homes the supermarket and the pharmacy are acids or bases for example aspirin is an acid acetylsalicylic acid and antacids are bases

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acid base equilibrium ionization that s what it s all about relative strengths a strong acid or base ionizes completely in aqueous solution 100 ionized or very darn close to it the equilibrium position lies far far to the right favors products since a strong acid base dissociates into the ions the concentration

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overview definitions of acids bases and salts identification of acids and bases description of the nature of proton in an aqueous solution explanation neutralization reactions explanation of how an acid base indicator works use of ph as a scale and discussion on

the importance of the ph value identification and preparation of salts

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introduction to acids bases and salts physical properties of acids physical properties of bases sources of various various acidsexamples of common saltsexamp

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acids and bases can be described using the arrhenius model acids produce  $H^+$  ions in aqueous solutions while bases produce  $OH^-$  ions we can identify acidic and basic solutions using their distinct and often contrasting properties some of which you are likely familiar with

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write chemical reactions between an arrhenius acid and an arrhenius base and between brønsted lowry acid and a brønsted lowry base there are three major classifications of substances known as acids or bases

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define an oxyacid and explain why many of these are very strong acids write an equation describing the amphoteric nature of zinc or aluminum hydroxide define an acid anhydride and write an equation describing its behavior explain how metal cations can give acidic solutions

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acid base basics chris ryan getty images acids produce protons or the  $H^+$  ion while bases accept protons or generate  $OH^-$  alternatively acids may be viewed as electron pair acceptors and bases as electron pair donors here are ways of defining acids and bases acids and bases and sample calculations acid base terms and definitions

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an acid is a substance which in aqueous solution produces hydroxonium ion  $H_3O^+$  or hydrogen ion  $H^+$  as the only positive ion also acids can be referred to as a proton donor classes of acids there are two classes of acids organic acids organic acids occur as natural products in plants and animal material inorganic acid

## **acid and base chart table of acids bases milliporesigma**

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use this acids and bases chart to find the relative strength of the most common acids and bases this acid base chart includes the  $K_a$  value for reference along with the chemical formula and the acid's conjugate base

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acid an acid is defined as a substance whose water solution tastes sour turns blue litmus red and neutralizes bases base a substance is called base if its aqueous solution tastes bitter turns red litmus blue or neutralizes acids salt salt is a neutral substance whose aqueous solution does not affect litmus recommended videos

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naming bases most bases are named just like ionic compounds the name of the cation is followed by the name of the anion for example koh is named potassium hydroxide nahco<sub>3</sub> is named sodium bicarbonate to write the formula of a base when given the name work backwards just like with ionic compounds

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acids and bases 2013 01 31 brian g cox this book seeks to enhance our understanding of acids and bases by reviewing and analysing their behaviour in non aqueous solvents the behaviour is related where possible to that in water but correlations and contrasts between solvents are also presented acids and bases 1968 russell s drago

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acids bases and salts chapter2 you have learnt in your previous classes that the sour and bitter tastes of food are due to acids and bases respectively present in them

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chemistry can be explained by variations of fundamental acid base concepts organic chemistry an acid base approach provides a framework for understanding the subject that goes beyond mere memorization

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