

EPUB FREE INTRODUCTION TO ACIDS BASES A WEBQUEST ANSWER KEY .PDF

ACIDS AND BASES CAN BE DESCRIBED USING THE ARRHENIUS MODEL. ACIDS PRODUCE H^+ IONS IN AQUEOUS SOLUTIONS WHILE BASES PRODUCE OH^- IONS. WE CAN IDENTIFY ACIDIC AND BASIC SOLUTIONS USING THEIR DISTINCT AND OFTEN CONTRASTING PROPERTIES SOME OF WHICH YOU ARE LIKELY FAMILIAR WITH. THE EARLIEST DEFINITION OF ACIDS AND BASES IS ARRHENIUS'S DEFINITION WHICH STATES THAT AN ACID IS A SUBSTANCE THAT FORMS HYDROGEN IONS H^+ WHEN DISSOLVED IN WATER AND A BASE IS A SUBSTANCE THAT FORMS HYDROXIDE IONS OH^- WHEN DISSOLVED IN WATER. SUGGEST SIMPLE TESTS YOU COULD CARRY OUT TO DETERMINE IF AN UNKNOWN SUBSTANCE IS AN ACID OR A BASE. STATE THE CHEMICAL DEFINITIONS OF AN ACID AND A BASE IN TERMS OF THEIR BEHAVIOR IN WATER. WRITE THE FORMULA OF THE SALT FORMED WHEN A GIVEN ACID AND BASE ARE COMBINED. ACID-BASE PROPERTIES OF SALTS. PH OF SALT SOLUTIONS. THIS UNIT IS PART OF THE CHEMISTRY LIBRARY. BROWSE VIDEOS, ARTICLES AND EXERCISES BY TOPIC. ACIDS AND BASES SOLUTIONS ARE CLASSIFIED AS ACIDIC OR BASIC BASED ON THEIR HYDROGEN ION CONCENTRATION RELATIVE TO PURE WATER. ACIDIC SOLUTIONS HAVE A HIGHER H^+ CONCENTRATION THAN WATER GREATER THAN $1 \times 10^{-7} M$ WHILE BASIC/ALKALINE SOLUTIONS HAVE A LOWER H^+ CONCENTRATION LESS THAN $1 \times 10^{-7} M$. THERE ARE THREE MAJOR CLASSIFICATIONS OF SUBSTANCES KNOWN AS ACIDS OR BASES. THE ARRHENIUS DEFINITION STATES THAT AN ACID PRODUCES H^+ IN SOLUTION AND A BASE PRODUCES OH^- . THIS THEORY WAS DEVELOPED BY SVANTE ARRHENIUS IN 1883. LATER TWO MORE SOPHISTICATED AND GENERAL THEORIES WERE PROPOSED. LEARN THE BASICS OF ACID-BASE CHEMISTRY. GET THE DEFINITIONS AND EXAMPLES OF ACIDS AND BASES AND LEARN ABOUT THE NEUTRALIZATION REACTION. ITS CHEMICAL FORMULA IS NH_3 . YOU CAN FIND IT IN MANY WINDOW-CLEANING PRODUCTS. NOT TO CONFUSE YOU BUT SCIENTISTS SOMETIMES USE ANOTHER SCHEME THE LEWIS SYSTEM TO DEFINE ACIDS AND BASES INSTEAD OF PROTONS. THIS LEWIS DEFINITION DESCRIBES WHAT MOLECULES DO WITH THEIR ELECTRONS. ACIDS AND BASES ARE COMMON SUBSTANCES FOUND IN MANY EVERY-DAY ITEMS FROM FRUIT JUICES AND SOFT DRINKS TO SOAP. IN THIS CHAPTER WE WILL EXAMINE THE PROPERTIES OF ACIDS AND BASES AND LEARN ABOUT THE CHEMICAL NATURE OF THESE IMPORTANT COMPOUNDS WE WILL COVER PH AND HOW TO CALCULATE THE PH OF A SOLUTION. A BASE IS A MOLECULE OR ION ABLE TO ACCEPT A HYDROGEN ION FROM AN ACID. ACIDIC SUBSTANCES ARE USUALLY IDENTIFIED BY THEIR SOUR TASTE. AN ACID IS BASICALLY A MOLECULE WHICH CAN DONATE AN H^+ ION AND CAN REMAIN ENERGETICALLY FAVOURABLE AFTER A LOSS OF H^+ . ACIDS ARE KNOWN TO TURN BLUE LITMUS RED. WHAT IS AN ACID OR A BASE? WHETHER A LIQUID IS AN ACID OR A BASE HAS TO DO WITH HYDROGEN IONS. ABBREVIATED WITH THE CHEMICAL SYMBOL H^+ IN WATER H_2O . A SMALL NUMBER OF THE MOLECULES DISSOCIATE/SPLIT UP. SOME OF THE WATER MOLECULES LOSE A HYDROGEN AND BECOME HYDROXIDE IONS OH^- . ACIDS AND BASES CAN BE STRONG OR WEAK DEPENDING ON THE EXTENT OF IONIZATION IN SOLUTION. MOST CHEMICAL REACTIONS REACH EQUILIBRIUM AT WHICH POINT THERE IS NO NET CHANGE. THE PH SCALE IS USED TO SUCCINCTLY COMMUNICATE THE ACIDITY OR BASICITY OF A SOLUTION. ACIDS AND BASES ARE IMPORTANT CLASSES OF CHEMICAL COMPOUNDS. THEY ARE PART OF THE FOODS AND BEVERAGES WE INGEST. THEY ARE PRESENT IN MEDICINES AND OTHER CONSUMER PRODUCTS AND THEY ARE PREVALENT IN THE WORLD AROUND US. IN THIS CHAPTER WE WILL FOCUS ON ACIDS AND BASES AND THEIR CHEMISTRY. UPDATED ON JULY 03 2019. LEARN ABOUT ACIDS, BASES AND PH INCLUDING DEFINITIONS AND CALCULATION. ACID-BASE BASICS. CHRIS RYAN GETTY IMAGES. ACIDS PRODUCE PROTONS OR THE H^+ ION WHILE BASES ACCEPT PROTONS OR GENERATE OH^- . ALTERNATIVELY ACIDS MAY BE VIEWED AS ELECTRON PAIR ACCEPTORS AND BASES AS ELECTRON PAIR DONORS. ARRHENIUS DEFINES ACIDS AND BASES BY THE DISSOCIATION PRODUCTS THAT ARE FORMED WHEN THE ACID OR BASE IS ADDED TO WATER WHILE BRONSTED AND LOWRY DEFINE ACIDS AND BASES BY THE REACTIONS THAT OCCUR WHEN BOTH ARE ADDED TOGETHER. UPDATED ON MAY 05 2019. THERE ARE SEVERAL METHODS OF DEFINING ACIDS AND BASES WHILE THESE DEFINITIONS DON'T CONTRADICT EACH OTHER THEY DO VARY IN HOW INCLUSIVE THEY ARE. THE MOST COMMON DEFINITIONS OF ACIDS AND BASES ARE ARRHENIUS ACIDS AND BASES, BRONSTED-LOWRY ACIDS AND BASES AND LEWIS ACIDS AND BASES. ACID-BASE REACTIONS ARE ESSENTIAL IN BOTH BIOCHEMISTRY AND INDUSTRIAL CHEMISTRY. MOREOVER MANY OF THE SUBSTANCES WE ENCOUNTER IN OUR HOMES, THE SUPERMARKET AND THE PHARMACY ARE ACIDS OR BASES. FOR EXAMPLE ASPIRIN IS AN ACID (ACETYLSALICYLIC ACID) AND ANTACIDS ARE BASES. ACID AND BASE CHART LISTS THE STRENGTH OF ACIDS AND BASES STRONGEST TO WEAKEST IN ORDER. SIMPLE TO USE LABORATORY REFERENCE CHART FOR SCIENTISTS, RESEARCHERS AND LAB TECHNICIANS. HYPERCALCEMIA THAT IS SEVERE DEFINED AS TOTAL CALCIUM OF 14 mg/dL OR GREATER, 3.5 mmol/L OR IONIZED CALCIUM OF 10 mg/dL OR GREATER, 2.5 mmol/L OR THAT DEVELOPS RAPIDLY OVER DAYS TO WEEKS CAN CAUSE NAUSEA, VOMITING, DEHYDRATION, CONFUSION, SOMNOLENCE AND COMA. OBSERVATIONS APPROXIMATELY 90% OF PEOPLE WITH HYPERCALCEMIA HAVE PRIMARY. THE FOLLOWING EQUATION IS USED FOR CALCULATING ACID AND BASE MOLARITY WHERE THE CONCENTRATION IS GIVEN IN wt/d , MW 10 MOLARITY WHERE WEIGHT/DENSITY OR SPECIFIC GRAVITY, MW MOLECULAR WEIGHT OR FORMULA WEIGHT. THE ABOVE EQUATION CAN THEN BE USED TO CALCULATE THE MOLARITY OF THE 70% WT NITRIC ACID.

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