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the molar mass of a substance is the mass of one mole of the substance a mole of any substance contains 6 022137 1023 particles or avogadro s number of particles the atomic mass is the mass of one atom 24 the atomic mass of a single isotope of an element is its relative mass compared to the mass of carbon 12

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1 what is the relative mass and charge of a neutron mass 1 and charge 1 mass 0 and charge 1 mass 1 and charge 0 2 an atom has 13 protons and 14 neutrons what is its mass

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chapter 5 electrons in atoms chapter test practice your results the correct answer for each question is indicated by a 1 the wavelength of a wave is measured

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atom quiz thoughtco Jul 22 2023

atom quiz all matter is made of atoms take this fun quiz to see how much you know about the atom paper boat creative getty images 1 the three basic components of an atom are although they can be broken further into subatomic particles the three main components of an atom are protons neutrons and electrons

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atomic structure practice khan academy google classroom what three particles make up an atom choose 1 answer protons neutrons and photons a protons neutrons and photons positrons neutrons and electrons b positrons neutrons and electrons protons electrons and positrons c protons electrons and positrons

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chapter 1 atoms 1 1 early ideas in atomic theory the ancient greeks proposed that matter consists of extremely small particles called atoms

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write the letter of the correct answer on the line at the left a n is an atom or group of atoms that has an electric charge 2 compound b ion c crystal d neutron covalent bond is called if the atoms equally share one or more pairs of electrons double b ionic c nonpolar d polar

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the quantum mechanical model of the atom is concerned with the probability of finding an electron in a certain position configuration for calcium 1s1 2s2 2p2 3s2 3p3 4s2 18 the maximum number of electrons that can occupy the third principle energy level study with quizlet and memorize flashcards containing terms like hertz atomic

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chemistry the interaction of atoms and molecules with each other is principally governed by the electrons which are bound most weakly to the nucleus in this chapter we learn about the electronic structure of atoms and how this is related to the properties of the atoms

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the force of attraction that holds two atoms together is called a an 12 most atoms are less likely to react with other atoms when they have valence electrons 13 when an atom loses an electron a an ion forms 14 when two atoms share two pairs of electrons a an forms 15

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two atoms of the same element must have the same number of protons c it is impossible to simultaneously know the precise position and velocity of a particle d electrons of atoms in their ground states enter energetically equivalent sets of orbitals singly before they pair up in any orbital of the set 7

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